

The Crystal Structures of Derivatives of Tetrodotoxin. II. Diacetylanhydrotetrodotoxin Hydroiodide

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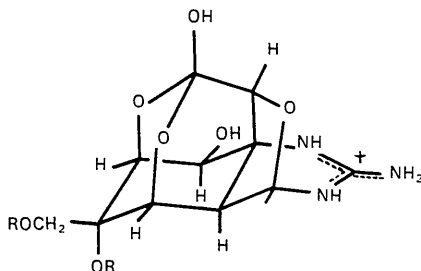
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The structure of diacetylanhydrotetrodotoxin, one of the minimum changed derivatives of the toxic substance tetrodotoxin, was determined by X-ray analysis of the hydroiodide. The crystals are monoclinic with four molecules in a unit cell of dimensions $a=19.34$, $b=14.63$, $c=7.18$ Å, $\beta=97.4^\circ$, with space group $P2_1$. There are two independent molecules in an asymmetric unit. The positions of iodine atoms were determined by three-dimensional Patterson methods. After several cycles of three-dimensional Fourier refinement, it was found that the two independent molecules have essentially the same structure,



($R = \text{CH}_3\text{CO}$), which involves an adamantane-like cage. Atomic coordinates and individual isotropic temperature factors (anisotropic for the iodine) were refined by the block-diagonal matrix least-squares method. The final R value decreased to 0.157. Intra- and intermolecular bond lengths are within the range of normal values.

Introduction

In a preliminary paper, Tsuda, Tamura, Tachikawa, Sakai, Amakasu, Kawamura & Ikuma (1963) proposed a lactam form (II) for the structure of tetrodotoxin. This was inferred from chemical considerations, taking account of an X-ray determination of the structure of tetrodonic acid hydrobromide (I) described in part I. Goto, Kishi, Takahashi & Hirata (1963) proposed independently the same structure for it on the basis of an X-ray analysis of bromoanhydrotetrodoic lactone hydrobromide (III) (Tomiie, Furusaki, Kasami, Yasuoka, Miyake, Haisa & Nitta, 1963). Further studies were required, however, since several other possible structures were not excluded definitely.

Diacetylanhydrotetrodotoxin (hereafter DAT) was prepared by alcoholysis of peracetyltetrodotoxin, and its hydroiodide (hereafter DATI) was obtained by careful control of the acidity of the mother liquor. Tetrodotoxin can be recovered from DAT or DATI by treating its alcoholic solution with a drop of aqueous hydrochloric acid (5%), but so far not from tetrodonic acid. This fact suggests that the chemical relationship

between tetrodotoxin and DAT (DATI) may be very close.

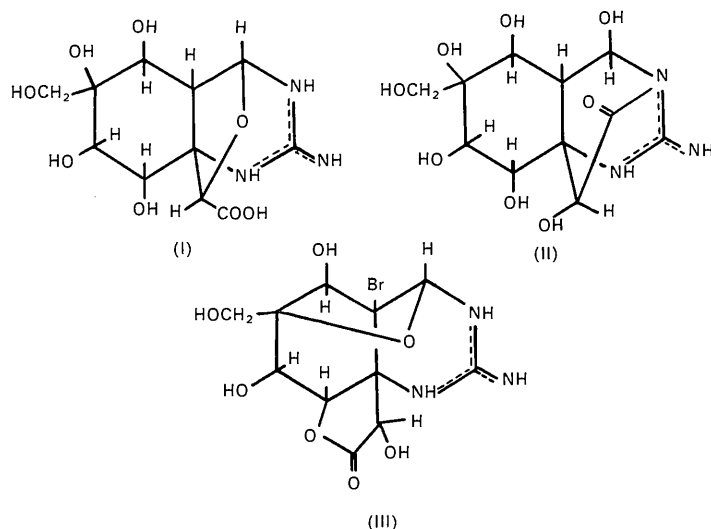
The present paper describes an X-ray determination of diacetylanhydrotetrodotoxin hydroiodide. On the basis of the structure of this compound, the structure of tetrodotoxin was established by some accompanying chemical work. The final chemical report has already been published by Tsuda, Ikuma, Kawamura, Tachikawa, Sakai, Tamura & Amakasu (1964). Meanwhile, the same structure has been presented independently by Goto, Kishi, Takahashi & Hirata (1964).

Experimental

Crystals of DATI, in the form of colourless plates, were grown by adding ether to a methanol solution containing hydroiodic acid. The unit-cell dimensions of the monoclinic crystals are:

$$a=19.34, b=14.63, c=7.18 \text{ \AA}, \beta=97.4^\circ.$$

The space group was uniquely determined to be $P2_1$ from the systematic absences of $0k0$ with odd k and the fact that this compound is a naturally occurring



product with asymmetric carbon atoms. The density measured by the flotation method in a mixture of carbon tetrachloride and tetrabromoethane was found to be 1.69 g.cm^{-3} . From the unit-cell dimensions and the density, the molecular weight was calculated to be 1026, assuming that the number of the molecules in the unit cell is two. In a later stage of the analysis, however, it was found that there are four molecules in the unit cell, that is, two crystallographically independent molecules in an asymmetric unit. The molecular weight was therefore taken to be 513.

Equi-inclination Weissenberg photographs were taken for the zero to 12th layers about the b axis with $\text{Cu K}\alpha$ radiation. To obtain scaling factors among various k layers, the photographs around the c axis were taken by the same method.

Intensities were measured by visual comparison with a standard intensity scale. Strong and weak reflexions were correlated by the multiple-film technique. The maximum value of $\sin \theta$ observed was 0.985, and the intensity data for 3004 reflexions were collected out of 4003 accessible hkl 's.

The sizes of crystals used were less than $0.1 \times 0.5 \times 0.8$ mm. The crystals gradually became yellow and then brown, and the intensities of reflexions in higher angles decreased slightly on exposure to X-rays for some 300 hours. Three crystals were used to collect the complete set of intensity data. Corrections for Lorentz and polarization factors were made in the usual way but those for absorption were neglected. Wilson's method gave an overall temperature factor of $B = 1.8 \text{ \AA}^2$. This unexpectedly low value of the temperature factor might be due to the effect of absorption.

Structure determination

Four remarkable peaks were found in the three-dimensional Patterson function. These were thought to correspond to the interatomic vectors due to two iodine atoms in the asymmetric unit. The y coordinate of one

of the iodine atoms could be put arbitrarily equal to zero. The two heavy atoms in the asymmetric unit eliminated the phase ambiguity inherent in space group $P2_1$ (Fridrichsons, Mathieson & Sutor, 1963).

The first three-dimensional Fourier synthesis, shown in Fig. 1, was calculated with the phases based on two iodine atoms. The number of peaks in this electron distribution map was much larger than that expected from the tentative chemical formula. Some of them should be ghost peaks due to the large scattering factor of the iodine atom. The clues for selecting peaks were the peak heights and their characteristic arrangement with reasonable interatomic distances. Examples of the latter are shown by connecting peaks in Fig. 2. Some thirty-six atoms were introduced into the structure factor calculation. In the second electron density distribution map, shown in Fig. 2, it was possible to connect many peaks reasonably at plausible bond distances. The structure factor calculation was made with fifty-three light atoms. The arrangement of significant peaks in the resulting Fourier map strongly suggested that there were two independent molecules in an asymmetric unit, although the detailed structures of the individual molecules were still ambiguous at this stage.

The $(F_o - F_c)$ synthesis was computed in the next cycle of refinement with the phases calculated from fifty-eight light atoms. Some of the atoms included in the earlier stage had to be excluded later. In the structure factor calculation so far, all light atoms were treated as nitrogen, with the isotropic overall temperature factor, $B = 1.8 \text{ \AA}^2$. At this stage, two systems of joined six-membered rings were recognized in one molecule in the asymmetric unit. It was very difficult to explain this grouping of peaks by the hypothetical lactam form (II). After several different trial models of the skeleton had been examined, it was concluded that the orthoester model could fit the constellation of peaks in the Fourier map. Such a constellation of peaks was also found in another molecule in the asymmetric unit, and this result strongly supported the reliability

of the analysis. Other functional groups were connected to the skeleton. Thus the two molecules so built up have essentially the same structure, as shown by formula (IV).

Besides these two molecules and two iodines, a significant peak O(78) supposed to be crystallization water was observed. The measurement of magnetic resonance spectra indicated that there should be one methyl alcohol or methoxyl group per two DAT molecules. In order to find out the position of this methyl group, the $(F_o - F_c)$ synthesis was calculated, where F_c included two DAT molecules, two iodine atoms and O(78). The resulting map showed the peaks which might be bonded with O(21), O(71) and O(78) at reasonable distances. However, if the peak near O(21) is assumed to be the carbon atom, it results in unreasonably close approaches to the atoms in the neighbouring acetyl groups at distances of 2.0, 2.1 and 2.3 Å. If the carbon is assumed to attach to O(71), there results also

an unreasonable contact with the iodine atom at a distance of 2.7 Å. Because of the absence of abnormal intermolecular contacts, the carbon atom was decided as bonding with O(78), forming crystallization methyl alcohol.

Some systematic disagreement between the observed and calculated structure factors was observed for the reflexions at low angles. This might have occurred from neglecting the correction for absorption. Plots of F_c/F_o for hkl with fixed k and l against $(\sin \theta/\lambda)^2$, an example of which is shown in Fig. 4, indicated a large anisotropy of the absorption effect. For convenience, we corrected the observed structure factors for this anisotropic effect by the use of the experimental curves of the type shown in Fig. 4.

Atomic coordinates and temperature factors were refined by the block-diagonal matrix least-squares method. The temperature factors were taken as isotropic for the light atoms and as anisotropic for iodine. In the first two cycles, about 1000 reflexions for $\sin \theta$

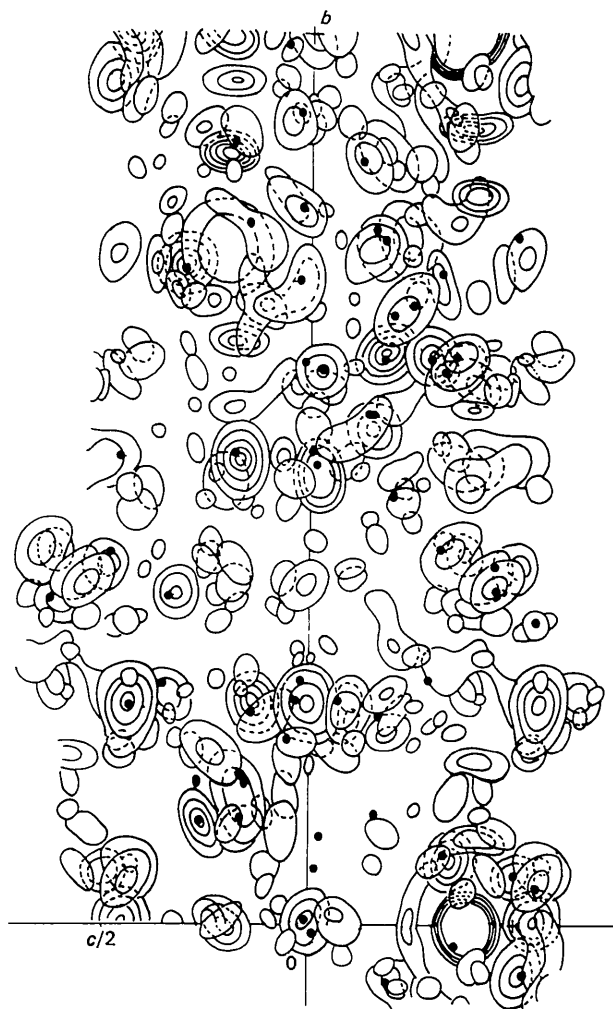


Fig. 1. The composite Fourier map projected along the a axis. Phases are based on two iodine atoms. Contour interval is arbitrary. Black circlets indicate the final atomic positions of two molecules in the asymmetric unit.

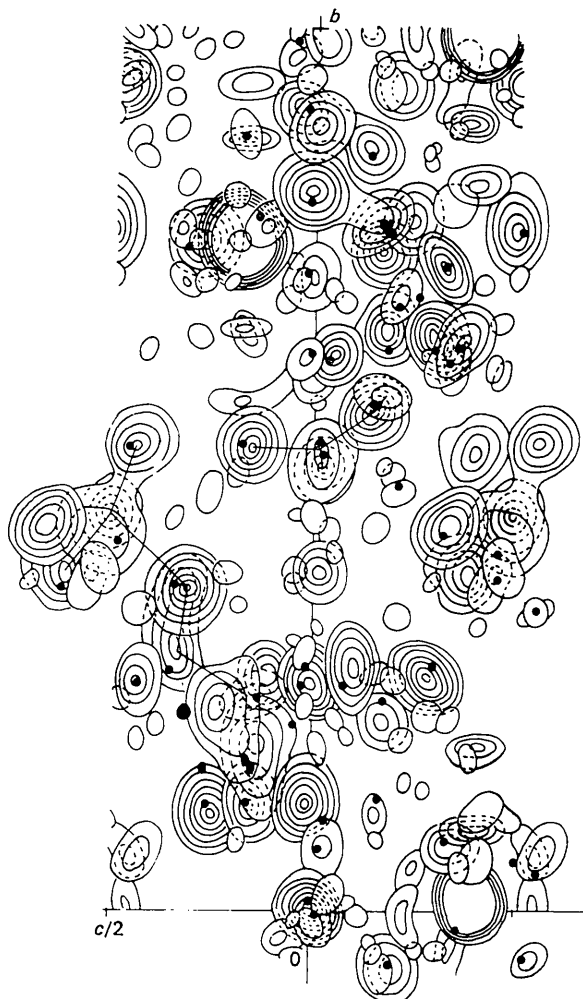


Fig. 2. The composite Fourier map projected along the a axis. Phases are based on two iodine and thirty-six light atoms. Contour interval is arbitrary. Black circlets indicate the final atomic positions of two molecules.

ranging between 0.22 and 0.42 were included. Three more cycles of refinement were made, using all the reflexions. The final *R* value dropped to 0.157. The final atomic coordinates and temperature factors for each atom are given in Table 1, and the observed and

Table 1. *The final atomic coordinates and temperature factors (Å²)*

For the iodine atoms, anisotropic temperature factors in the form of $\exp \{- (h^2 B_{11} + k^2 B_{22} + l^2 B_{33} + hk B_{12} + kl B_{23} + lh B_{31}) \}$ were applied.



Fig. 3. The final composite Fourier map projected along the *a* axis, contours being drawn at intervals of 1 e.Å⁻³. Contours for the iodine atoms are omitted and their positions are shown with black circles.

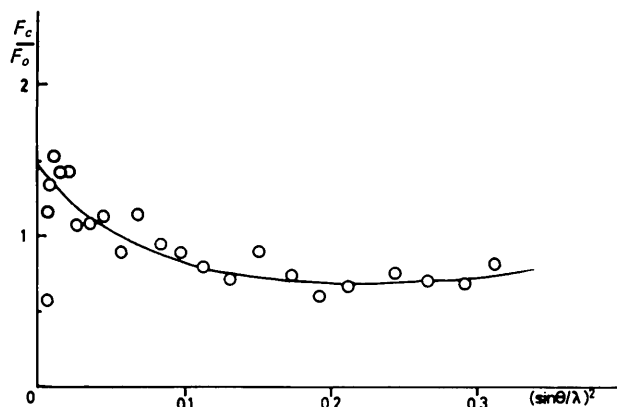


Fig. 4. F_c/F_o value against $(\sin \theta/\lambda)^2$.

Atom	<i>x/a</i>	<i>y/b</i>	<i>z/c</i>	<i>B</i>
N(1)	0.355	0.151	-0.122	1.2
C(2)	0.362	0.107	0.033	0.7
N(3)	0.406	0.137	0.182	2.0
C(4)	0.445	0.221	0.155	1.4
C(5)	0.418	0.376	0.013	2.5
C(6)	0.377	0.437	-0.157	0.7
C(7)	0.381	0.375	-0.329	1.5
C(8)	0.346	0.280	-0.327	1.5
C(9)	0.387	0.235	-0.156	1.3
C(10)	0.392	0.284	0.029	0.9
N(11)	0.332	0.025	0.056	2.8
O(12)	0.492	0.378	-0.033	1.5
O(13)	0.304	0.436	-0.141	1.9
C(14)	0.413	0.524	-0.182	3.1
O(15)	0.364	0.585	-0.311	2.7
O(16)	0.454	0.365	-0.354	1.9
O(17)	0.346	0.224	-0.493	1.6
O(18)	0.493	0.203	0.024	1.7
C(19)	0.465	0.231	-0.160	1.6
C(20)	0.495	0.319	-0.204	1.2
O(21)	0.561	0.313	-0.237	3.4
C(22)	0.287	0.491	0.006	2.6
O(23)	0.321	0.495	0.165	3.0
C(24)	0.203	0.501	-0.020	3.2
C(25)	0.395	0.621	-0.445	1.6
C(26)	0.343	0.674	-0.559	2.9
O(27)	0.457	0.609	-0.464	3.5
N(51)	0.013	0.931	-0.329	0.5
C(52)	0.026	1.015	-0.355	1.7
N(53)	0.051	1.069	-0.198	0.9
C(54)	0.055	1.020	-0.005	1.0
C(55)	0.114	0.883	0.158	0.0
C(56)	0.128	0.776	0.146	0.7
C(57)	0.061	0.736	0.025	1.0
C(58)	0.042	0.775	-0.168	0.0
C(59)	0.029	0.887	-0.152	0.0
C(60)	0.090	0.932	-0.028	0.7
N(61)	0.011	1.055	-0.523	1.8
O(62)	0.052	0.887	0.264	0.7
O(63)	0.180	0.761	0.016	2.1
C(64)	0.146	0.736	0.338	3.0
O(65)	0.157	0.642	0.329	1.0
O(66)	0.001	0.756	0.134	0.9
O(67)	-0.013	0.736	-0.274	1.1
O(68)	-0.014	0.998	0.026	0.4
C(69)	-0.026	0.906	-0.019	0.8
C(70)	-0.009	0.852	0.175	2.0
O(71)	-0.057	0.860	0.295	1.2
C(72)	0.241	0.804	0.058	2.2
O(73)	0.267	0.858	0.166	3.8
C(74)	0.287	0.761	-0.086	2.6
C(75)	0.162	0.598	0.483	0.7
C(76)	0.165	0.497	0.462	2.5
O(77)	0.162	0.644	0.634	2.0
O(78)	0.390	0.961	-0.553	3.2
C(79)	0.433	0.949	-0.376	3.5
I(81)	0.2253	0.0000	-0.3941	*
I(82)	0.1902	0.2433	0.1903	†

*B*₁₁ *B*₂₂ *B*₃₃ *B*₁₂ *B*₂₃ *B*₃₁ (× 10⁵)

* I(81) 121 123 543 -50 -306 51
 † I(82) 94 116 1195 97 296 13

Table 2. Observed and calculated structure factors (x 10)

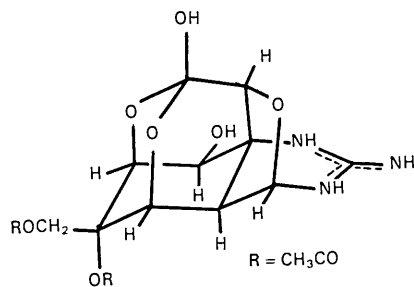
Table with multiple columns (h, k, l, Fo, Fc) and rows representing different data points and reflections. The table is organized into several vertical sections, likely representing different crystallographic planes or reflections. Each section contains observed (Fo) and calculated (Fc) structure factor values, along with some numerical indices or identifiers.

Table 2 (cont.)

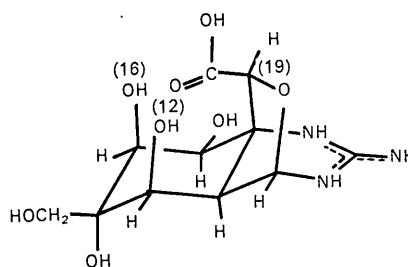
A large table with multiple columns and rows of numerical data, organized into several sections with headers like 'h', 'Po', 'Pc', 'h', 'Po', 'Pc', etc. The data includes various integers and some floating-point numbers, with some rows containing specific identifiers or groupings.

Table 2 (cont.)

h	h ₀	F _c	h	h ₀	F _c	h	h ₀	F _c	h	h ₀	F _c	h	h ₀	F _c	h	h ₀	F _c	h	h ₀	F _c	h	h ₀	F _c			
12	2		16	121	103	14	150	119	11	130	73	10	0	132	4	212	204	16	168	156	12	209	226	9	0	141
1	225	206	12	1		15	272	336	12	0	82	11	332	365	5	171	100	17	134	78	13	220	244	10	0	129
2	201	177				16	0	17	13	60	95	12	253	258	6	200	187	18	0	18	14	0	84	11	142	62
3	450	512	1	181	186	17	231	246	14	21	113	13	153	53	7	122	138	2	156	80	15	238	202	12	0	156
4	284	269	2	182	71	12	0		15	0	63	14	336	295	8	173	184	3	389	390	16	172	61	13	0	62
5	262	237	3	210	170				16	270	275	10	0	81	9	191	113	4	293	274	12	-4		14	0	59
6	372	466	4	154	149	1	180	298	0	121	70	11	0	35	11	0	35	5	172	161	0	112	143			
7	135	25	5	167	175	2	312	471	1	187	155	12	113	94	12	113	94	7	229	156	1	0	82			
8	201	272	6	189	163	3	280	283	2	233	230	13	130	121	13	130	121	8	303	386	2	209	111	0	319	340
9	293	287	7	217	225	4	317	298	3	186	153	14	138	126	15	124	95	9	210	213	3	233	198	1	196	141
10	277	313	8	293	280	5	436	432	4	185	150	15	124	95	16	148	230	10	192	250	4	264	250	2	241	267
11	214	203	9	182	134	6	237	158	5	172	170	16	148	230	17	0	56	12	0	43	5	143	45	3	192	207
12	0	71	10	269	276	7	263	261	6	156	213	17	0	56	12	0	56	12	0	43	6	0	69	4	240	179
13	191	172	11	264	227	8	152	175	7	306	453	18	74	149	18	74	149	13	198	240	8	216	201	6	222	54
14	0	75	12	269	309	9	235	233	8	0	75	19	165	192	19	165	192	14	139	113	9	199	187	7	277	318
15	168	193	13	187	209	10	174	180	9	291	359	2	224	236	2	224	236	15	156	184	11	134	161	8	215	167



(IV)



(V)

calculated structure factors are listed in Table 2. Computations were done on an IBM 7090 computer with programs written by the authors.

Discussion

Fig. 3 shows the superimposed sections of the electron density distribution cut through the atomic centres of DATI and methyl alcohol. It is conceived that the general structure of the DAT molecule is now established, except its absolute configuration. The fact that essentially the same structure has been deduced from the two independent molecules in the asymmetric unit is additional evidence of the reliability of the present analysis. In the present paper, except in Figs. 1, 2 and 3, the molecules are drawn with the correct absolute configuration, by taking account of the result by Tomiie, Furusaki, Kasami, Yasuoka, Miyake, Haisa & Nitta (1964) who have determined the absolute configuration of bromoanhydrotetrodoic lactone hydrobromide, another derivative of tetrodotoxin.

The present compound and tetrodonic acid have the same perhydroquinazoline skeleton. The conformation of substituent atoms to the cyclohexane ring is the same for these two compounds; O(17) and C(14) are in equatorial positions and O(13) is in axial position. The structural relationship between DAT and tetrodonic acid may be described as follows: Let us start from the hypothetical epitetrodonic acid (V), in which the configuration at C(19) is antipodal to that in tetrodonic acid. The ortho-ester may, then, be formed by combining the acid carbonyl with the hydroxyl groups, O(12) and O(16). The skeletal structure of DAT so built up is similar to that of adamantane, which is a cage molecule consisting of four six-membered rings in the chair form.

Intramolecular bond lengths and angles in the two independent molecules are shown in Fig. 5 and Table 3 respectively. The estimated standard deviations in the atomic coordinates are approximately 0.004 Å for the iodine and 0.07 Å for the light atoms. Hence, the standard deviation of the bond lengths between the light

Table 3. Bond angles

C(2)—N(1)—C(9)	126°	C(52)—N(51)—C(59)	123°
N(1)—C(2)—N(3)	121	N(51)—C(52)—N(53)	119
N(1)—C(2)—N(11)	124	N(51)—C(52)—N(61)	122
N(3)—C(2)—N(11)	115	N(53)—C(52)—N(61)	119
C(2)—N(3)—C(4)	117	C(52)—N(53)—C(54)	115
N(3)—C(4)—C(10)	105	N(53)—C(54)—C(60)	106
N(3)—C(4)—O(18)	109	N(53)—C(54)—O(68)	108
C(10)—C(4)—O(18)	99	C(60)—C(54)—O(68)	106
C(6)—C(5)—C(10)	115	C(56)—C(55)—C(60)	117
C(6)—C(5)—O(12)	102	C(56)—C(55)—O(62)	103
C(10)—C(5)—O(12)	113	C(60)—C(55)—O(62)	105
C(5)—C(6)—C(7)	102	C(55)—C(56)—C(57)	105
C(5)—C(6)—O(13)	109	C(55)—C(56)—O(63)	108
C(5)—C(6)—C(14)	112	C(55)—C(56)—C(64)	116
C(7)—C(6)—O(13)	102	C(57)—C(56)—O(63)	100
C(7)—C(6)—C(14)	110	C(57)—C(56)—C(64)	116
O(13)—C(6)—C(14)	120	O(63)—C(56)—C(64)	115
C(6)—C(7)—C(8)	117	C(56)—C(57)—C(58)	117
C(6)—C(7)—O(16)	108	C(56)—C(57)—O(66)	110
C(8)—C(7)—O(16)	110	C(58)—C(57)—O(66)	107
C(7)—C(8)—C(9)	102	C(57)—C(58)—C(59)	109
C(7)—C(8)—O(17)	117	C(57)—C(58)—O(67)	116
C(9)—C(8)—O(17)	111	C(59)—C(58)—O(67)	110
N(1)—C(9)—C(8)	109	N(51)—C(59)—C(58)	114
N(1)—C(9)—C(10)	104	N(51)—C(59)—C(60)	113
N(1)—C(9)—C(19)	115	N(51)—C(59)—C(69)	113
C(8)—C(9)—C(10)	118	C(58)—C(59)—C(60)	111
C(8)—C(9)—C(19)	115	C(58)—C(59)—C(69)	110
C(10)—C(9)—C(19)	95	C(60)—C(59)—C(69)	96
C(4)—C(10)—C(5)	113	C(54)—C(60)—C(55)	114
C(4)—C(10)—C(9)	101	C(54)—C(60)—C(59)	96
C(5)—C(10)—C(9)	111	C(55)—C(60)—C(59)	115
C(5)—O(12)—C(20)	108	C(55)—O(62)—C(70)	116
C(6)—O(13)—C(22)	113	C(56)—O(63)—C(72)	116

Table 3 (cont.)

C(6)—C(14)—O(15)	108	C(56)—C(64)—O(65)	111
C(14)—O(15)—C(25)	112	C(64)—O(65)—C(75)	117
C(7)—O(16)—C(20)	114	C(57)—O(66)—C(70)	115
C(4)—O(18)—C(19)	111	C(54)—O(68)—C(69)	108
C(9)—C(19)—O(18)	105	C(59)—C(69)—O(68)	102
C(9)—C(19)—C(20)	113	C(59)—C(69)—C(70)	112
O(18)—C(19)—C(20)	110	O(68)—C(69)—C(70)	105
O(12)—C(20)—O(16)	106	O(62)—C(70)—O(66)	109
O(12)—C(20)—C(19)	106	O(62)—C(70)—C(69)	107
O(12)—C(20)—O(21)	109	O(62)—C(70)—O(71)	107
O(16)—C(20)—C(19)	112	O(66)—C(70)—C(69)	109
O(16)—C(20)—O(21)	111	O(66)—C(70)—O(71)	109
C(19)—C(20)—O(21)	113	C(69)—C(70)—O(71)	116
O(13)—C(22)—O(23)	125	O(63)—C(72)—O(73)	139
O(13)—C(22)—C(24)	107	O(63)—C(72)—C(74)	103
O(23)—C(22)—C(24)	120	O(73)—C(72)—C(74)	118
O(15)—C(25)—C(26)	106	O(65)—C(75)—C(76)	114
O(15)—C(25)—O(27)	125	O(65)—C(75)—O(77)	118
C(26)—C(25)—O(27)	129	C(76)—C(75)—O(77)	128

atoms is about 0.1 Å, and that of the bond angles is about 6°. None of the bond lengths and angles deviate significantly from their normal values. The average values of the bond lengths are 1.53 Å for C—C, 1.33 Å for C—N (in the guanidine group), 1.46 Å for C—N, 1.42 Å for C—O and 1.23 Å for C=O. It should be pointed out that the corresponding bond lengths and angles in the two independent molecules are in good agreement as seen from Fig. 5 and Table 3.

The crystal structure viewed along the *c* axis is illustrated in Fig. 6, where anticipated hydrogen bonds are shown by the broken lines. The hydrogen bond lengths and short intermolecular contacts are listed in Table 4. The donors of the hydrogen bonds in Table 4 are the atoms at *x*, *y*, *z*. There are six NH...O, three OH...N and one OH...O hydrogen bond per asymmetric unit, and these hydrogen bonds connect the

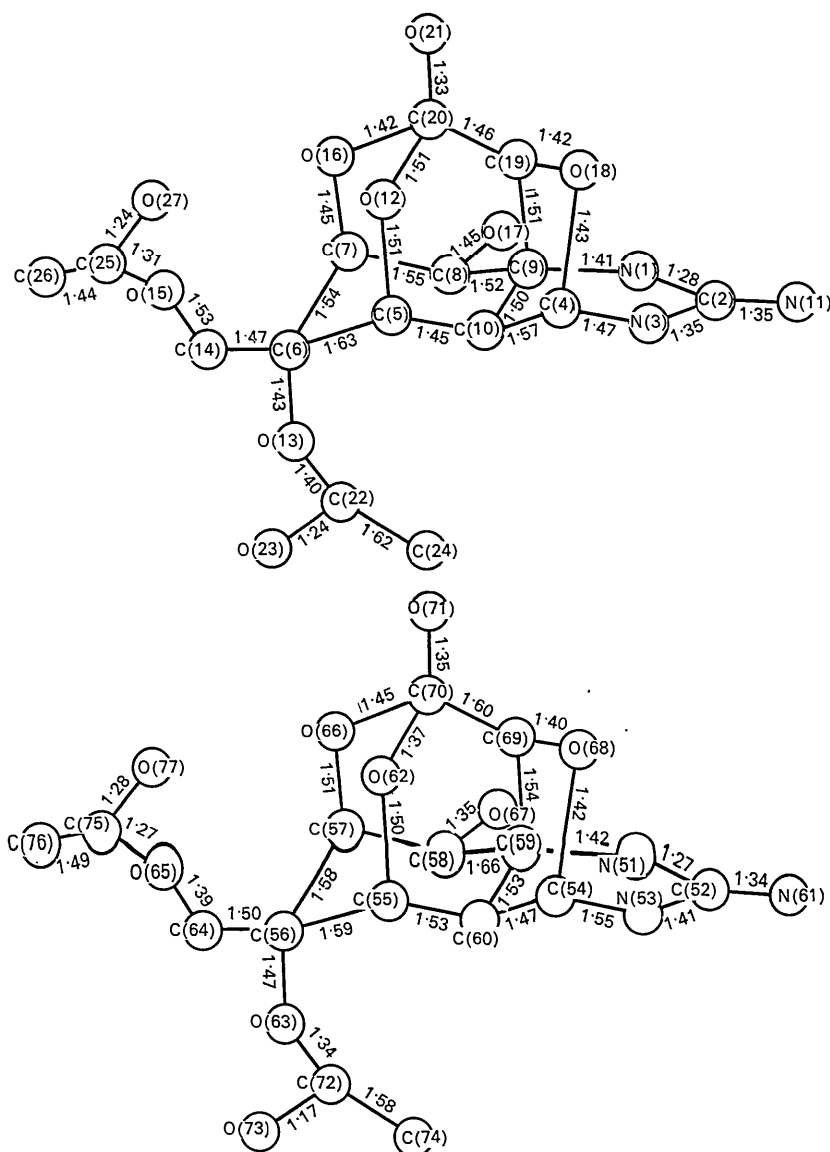
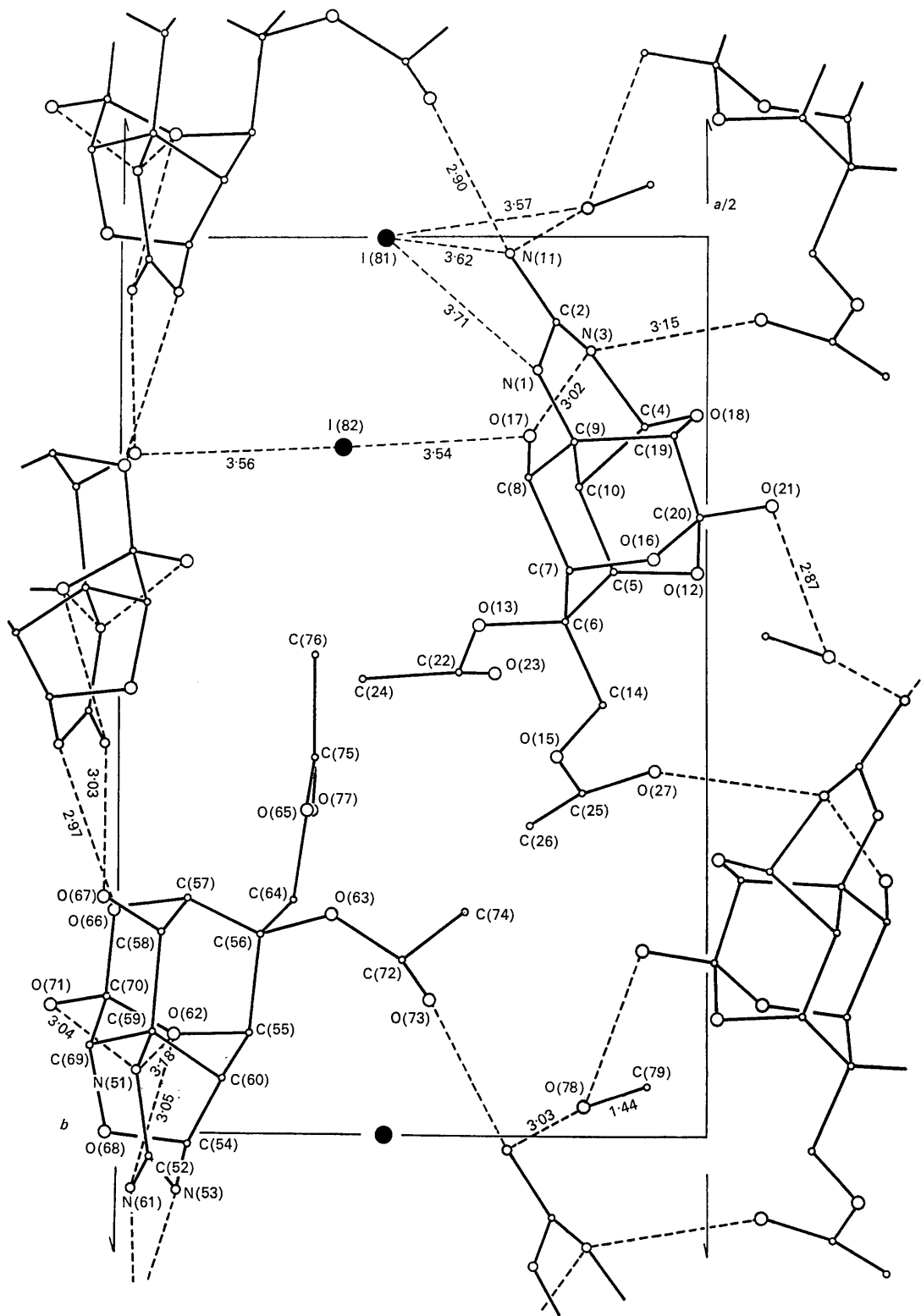


Fig. 5. Bond lengths (Å).



c downward
 Fig. 6. Crystal structure viewed along the *c* axis.

DAT and methyl alcohol molecules to form an infinite three-dimensional network. The iodine atoms are in the interstices of the three-dimensional network and short contacts are found for I(81)···N(1), I(81)···N(11), I(82)···O(17) and I(82)···O(67). It may be suggested that the iodine atoms interact with the hydrogen atoms bonded with these nitrogen or oxygen atoms.

Table 4. Possible hydrogen bonds and short intermolecular distances

From atom at x, y, z	To atom	At position	Distance
(a) Possible hydrogen bonds			
N(3)	O(27)	11	3.15 Å
N(11)	O(73)	4	2.90
O(17)	N(3)	2	3.02
O(21)	O(78)	12	2.87
N(51)	O(62)	2	3.18
N(53)	O(66)	9	2.97
N(61)	O(62)	2	3.05
N(61)	O(67)	10	3.03
O(71)	N(51)	3	3.04
O(78)	N(11)	5	3.03
(b) Short intermolecular contacts			
I(81)	N(1)	1	3.71
I(81)	N(11)	1	3.62
I(81)	O(78)	4	3.57
I(81)	N(53)	4	3.95
I(81)	C(52)	4	3.91
I(81)	O(73)	6	3.95
I(82)	O(17)	3	3.54
I(82)	O(67)	7	3.56
O(15)	C(74)	1	3.48
C(24)	O(65)	1	3.45
C(24)	C(76)	1	3.63
C(26)	C(74)	1	3.91
O(17)	C(4)	2	3.36
C(24)	C(75)	2	3.83
C(24)	C(76)	2	3.70
C(24)	O(77)	2	3.27
C(25)	O(23)	2	3.50
C(26)	O(23)	2	3.28
C(26)	C(64)	2	3.89
C(26)	C(72)	2	3.70
C(26)	O(73)	2	3.54
C(26)	C(75)	2	3.72
C(52)	O(62)	2	3.41
C(58)	O(77)	2	3.45
N(61)	O(68)	2	3.32
N(61)	O(71)	2	3.34
O(63)	O(77)	2	3.21
O(67)	O(71)	2	3.59
C(74)	O(77)	2	3.40
O(78)	O(73)	2	3.28
O(78)	N(3)	5	3.24
C(24)	C(69)	7	3.74
C(57)	N(53)	7	3.59
C(57)	C(54)	7	3.87
C(75)	C(52)	7	3.83
C(75)	N(61)	7	3.45
C(76)	N(51)	7	3.58
C(76)	C(52)	7	3.68
C(76)	N(61)	7	3.59

Table 4 (cont.)

From atom at x, y, z	To atom	At position	Distance				
O(77)	N(61)	7	3.58				
C(76)	O(71)	8	3.52				
C(4)	C(14)	11	3.97				
C(4)	C(25)	11	3.79				
C(4)	O(27)	11	3.18				
C(5)	C(79)	11	3.78				
O(12)	C(79)	11	3.27				
O(18)	C(14)	11	3.30				
O(18)	C(27)	11	3.47				
O(21)	O(73)	11	3.36				
O(21)	C(74)	11	3.58				
O(16)	O(78)	12	3.47				
O(16)	C(79)	12	3.34				
C(20)	C(79)	12	3.97				
O(21)	C(26)	12	3.23				
O(21)	C(79)	12	3.43				
O(27)	C(79)	12	3.46				
Position 1	x	y	z	7	$-x$	$-\frac{1}{2}+y$	$-z$
2	x	y	$-1+z$	8	$-x$	$-\frac{1}{2}+y$	$1-z$
3	x	y	$1+z$	9	$-x$	$\frac{1}{2}+y$	$-z$
4	x	$-1+y$	z	10	$-x$	$\frac{1}{2}+y$	$-1-z$
5	x	$1+y$	$-1+z$	11	$1-x$	$-\frac{1}{2}+y$	$-z$
6	x	$-1+y$	$-1+z$	12	$1-x$	$-\frac{1}{2}+y$	$-1-z$

All distances less than 4.0 Å between the reference molecules at x, y, z and the neighbouring molecules were calculated. Only those less than 4.0 Å for C···C and C···I pairs and those less than 3.6 Å for the other pairs are listed.

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References

- FRIDRICHSONS, J., MATHIESON, A. M. & SUTOR, D. J. (1963). *Acta Cryst.* **16**, 1075.
- GOTO, T., KISHI, Y., TAKAHASHI, S. & HIRATA, Y. (1963). *Tetrahedron Letters*, **30**, 2115.
- GOTO, T., KISHI, Y., TAKAHASHI, S. & HIRATA, Y. (1964). *Tetrahedron Letters*, **14**, 779.
- TOMIIE, Y., FURUSAKI, A., KASAMI, K., YASUOKA, N., MIYAKE, K., HAISA, M. & NITTA, I. (1963). *Tetrahedron Letters*, **30**, 2102.
- TOMIIE, Y., FURUSAKI, A., KASAMI, K., YASUOKA, N., MIYAKE, K., HAISA, M. & NITTA, I. (1964). IUPAC Symposium on the Chemistry of Natural Products (Kyoto).
- TSUDA, K., IKUMA, S., KAWAMURA, M., TACHIKAWA, R., SAKAI, K., TAMURA, C. & AMAKASU, O. (1964). *Chem. Pharm. Bull. (Tokyo)*, **12**, 1357.
- TSUDA, K., TAMURA, C., TACHIKAWA, R., SAKAI, K., AMAKASU, O., KAWAMURA, M. & IKUMA, S. (1963). *Chem. Pharm. Bull. (Tokyo)*, **11**, 1473.