

The Crystal Structures of Derivatives of Tetrodotoxin. II. Diacetylanhydrotetrodotoxin Hydroiodide

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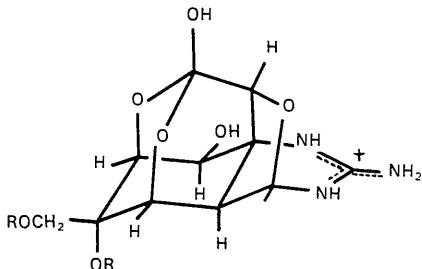
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The structure of diacetylanhydrotetrodotoxin, one of the minimum changed derivatives of the toxic substance tetrodotoxin, was determined by X-ray analysis of the hydroiodide. The crystals are monoclinic with four molecules in a unit cell of dimensions $a = 19.34$, $b = 14.63$, $c = 7.18 \text{ \AA}$, $\beta = 97.4^\circ$, with space group $P2_1$. There are two independent molecules in an asymmetric unit. The positions of iodine atoms were determined by three-dimensional Patterson methods. After several cycles of three-dimensional Fourier refinement, it was found that the two independent molecules have essentially the same structure,



($R = \text{CH}_3\text{CO}$), which involves an adamantane-like cage. Atomic coordinates and individual isotropic temperature factors (anisotropic for the iodine) were refined by the block-diagonal matrix least-squares method. The final R value decreased to 0.157. Intra- and intermolecular bond lengths are within the range of normal values.

Introduction

In a preliminary paper, Tsuda, Tamura, Tachikawa, Sakai, Amakasu, Kawamura & Ikuma (1963) proposed a lactam form (II) for the structure of tetrodotoxin. This was inferred from chemical considerations, taking account of an X-ray determination of the structure of tetrodonic acid hydrobromide (I) described in part I. Goto, Kishi, Takahashi & Hirata (1963) proposed independently the same structure for it on the basis of an X-ray analysis of bromoanhydrotetrodoic lactone hydrobromide (III) (Tomiiie, Furusaki, Kasami, Yasuoka, Miyake, Haisa & Nitta, 1963). Further studies were required, however, since several other possible structures were not excluded definitely.

Diacetylanhydrotetrodotoxin (hereafter DAT) was prepared by alcoholysis of peracetyletetrodotoxin, and its hydroiodide (hereafter DATI) was obtained by careful control of the acidity of the mother liquor. Tetrodotoxin can be recovered from DAT or DATI by treating its alcoholic solution with a drop of aqueous hydrochloric acid (5%), but so far not from tetrodonic acid. This fact suggests that the chemical relationship

between tetrodotoxin and DAT (DATI) may be very close.

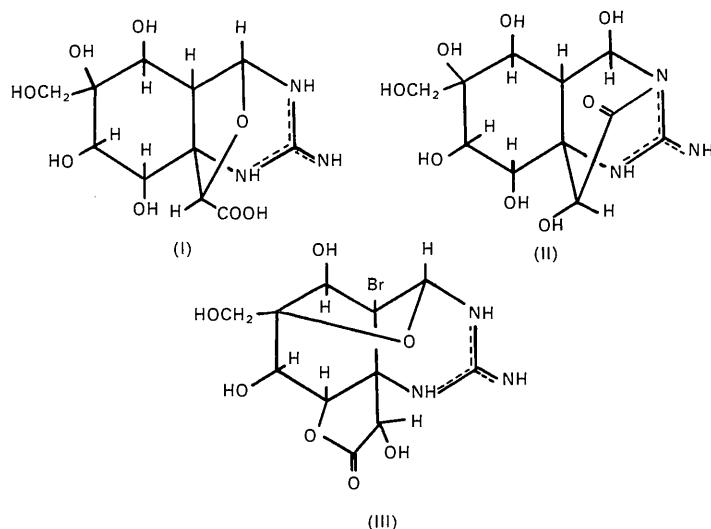
The present paper describes an X-ray determination of diacetylanhydrotetrodotoxin hydroiodide. On the basis of the structure of this compound, the structure of tetrodotoxin was established by some accompanying chemical work. The final chemical report has already been published by Tsuda, Ikuma, Kawamura, Tachikawa, Sakai, Tamura & Amakasu (1964). Meanwhile, the same structure has been presented independently by Goto, Kishi, Takahashi & Hirata (1964).

Experimental

Crystals of DATI, in the form of colourless plates, were grown by adding ether to a methanol solution containing hydroiodic acid. The unit-cell dimensions of the monoclinic crystals are:

$$a = 19.34, b = 14.63, c = 7.18 \text{ \AA}, \beta = 97.4^\circ.$$

The space group was uniquely determined to be $P2_1$ from the systematic absences of $0k0$ with odd k and the fact that this compound is a naturally occurring



product with asymmetric carbon atoms. The density measured by the flotation method in a mixture of carbon tetrachloride and tetrabromomethane was found to be 1.69 g.cm^{-3} . From the unit-cell dimensions and the density, the molecular weight was calculated to be 1026, assuming that the number of the molecules in the unit cell is two. In a later stage of the analysis, however, it was found that there are four molecules in the unit cell, that is, two crystallographically independent molecules in an asymmetric unit. The molecular weight was therefore taken to be 513.

Equi-inclination Weissenberg photographs were taken for the zero to 12th layers about the *b* axis with Cu $K\alpha$ radiation. To obtain scaling factors among various *k* layers, the photographs around the *c* axis were taken by the same method.

Intensities were measured by visual comparison with a standard intensity scale. Strong and weak reflexions were correlated by the multiple-film technique. The maximum value of $\sin \theta$ observed was 0.985, and the intensity data for 3004 reflexions were collected out of 4003 accessible *hkl*'s.

The sizes of crystals used were less than $0.1 \times 0.5 \times 0.8 \text{ mm}$. The crystals gradually became yellow and then brown, and the intensities of reflexions in higher angles decreased slightly on exposure to X-rays for some 300 hours. Three crystals were used to collect the complete set of intensity data. Corrections for Lorentz and polarization factors were made in the usual way but those for absorption were neglected. Wilson's method gave an overall temperature factor of $B = 1.8 \text{ \AA}^2$. This unexpectedly low value of the temperature factor might be due to the effect of absorption.

Structure determination

Four remarkable peaks were found in the three-dimensional Patterson function. These were thought to correspond to the interatomic vectors due to two iodine atoms in the asymmetric unit. The *y* coordinate of one

of the iodine atoms could be put arbitrarily equal to zero. The two heavy atoms in the asymmetric unit eliminated the phase ambiguity inherent in space group $P2_1$ (Fridrichsons, Mathieson & Sutor, 1963).

The first three-dimensional Fourier synthesis, shown in Fig. 1, was calculated with the phases based on two iodine atoms. The number of peaks in this electron distribution map was much larger than that expected from the tentative chemical formula. Some of them should be ghost peaks due to the large scattering factor of the iodine atom. The clues for selecting peaks were the peak heights and their characteristic arrangement with reasonable interatomic distances. Examples of the latter are shown by connecting peaks in Fig. 2. Some thirty-six atoms were introduced into the structure factor calculation. In the second electron density distribution map, shown in Fig. 2, it was possible to connect many peaks reasonably at plausible bond distances. The structure factor calculation was made with fifty-three light atoms. The arrangement of significant peaks in the resulting Fourier map strongly suggested that there were two independent molecules in an asymmetric unit, although the detailed structures of the individual molecules were still ambiguous at this stage.

The $(F_o - F_c)$ synthesis was computed in the next cycle of refinement with the phases calculated from fifty-eight light atoms. Some of the atoms included in the earlier stage had to be excluded later. In the structure factor calculation so far, all light atoms were treated as nitrogen, with the isotropic overall temperature factor, $B = 1.8 \text{ \AA}^2$. At this stage, two systems of joined six-membered rings were recognized in one molecule in the asymmetric unit. It was very difficult to explain this grouping of peaks by the hypothetical lactam form (II). After several different trial models of the skeleton had been examined, it was concluded that the orthoester model could fit the constellation of peaks in the Fourier map. Such a constellation of peaks was also found in another molecule in the asymmetric unit, and this result strongly supported the reliability

of the analysis. Other functional groups were connected to the skeleton. Thus the two molecules so built up have essentially the same structure, as shown by formula (IV).

Besides these two molecules and two iodines, a significant peak O(78) supposed to be crystallization water was observed. The measurement of magnetic resonance spectra indicated that there should be one methyl alcohol or methoxyl group per two DAT molecules. In order to find out the position of this methyl group, the $(F_o - F_c)$ synthesis was calculated, where F_c included two DAT molecules, two iodine atoms and O(78). The resulting map showed the peaks which might be bonded with O(21), O(71) and O(78) at reasonable distances. However, if the peak near O(21) is assumed to be the carbon atom, it results in unreasonably close approaches to the atoms in the neighbouring acetyl groups at distances of 2.0, 2.1 and 2.3 Å. If the carbon is assumed to attach to O(71), there results also

an unreasonable contact with the iodine atom at a distance of 2.7 Å. Because of the absence of abnormal intermolecular contacts, the carbon atom was decided as bonding with O(78), forming crystallization methyl alcohol.

Some systematic disagreement between the observed and calculated structure factors was observed for the reflexions at low angles. This might have occurred from neglecting the correction for absorption. Plots of F_c/F_o for hkl with fixed k and l against $(\sin \theta/\lambda)^2$, an example of which is shown in Fig. 4, indicated a large anisotropy of the absorption effect. For convenience, we corrected the observed structure factors for this anisotropic effect by the use of the experimental curves of the type shown in Fig. 4.

Atomic coordinates and temperature factors were refined by the block-diagonal matrix least-squares method. The temperature factors were taken as isotropic for the light atoms and as anisotropic for iodine. In the first two cycles, about 1000 reflexions for $\sin \theta$

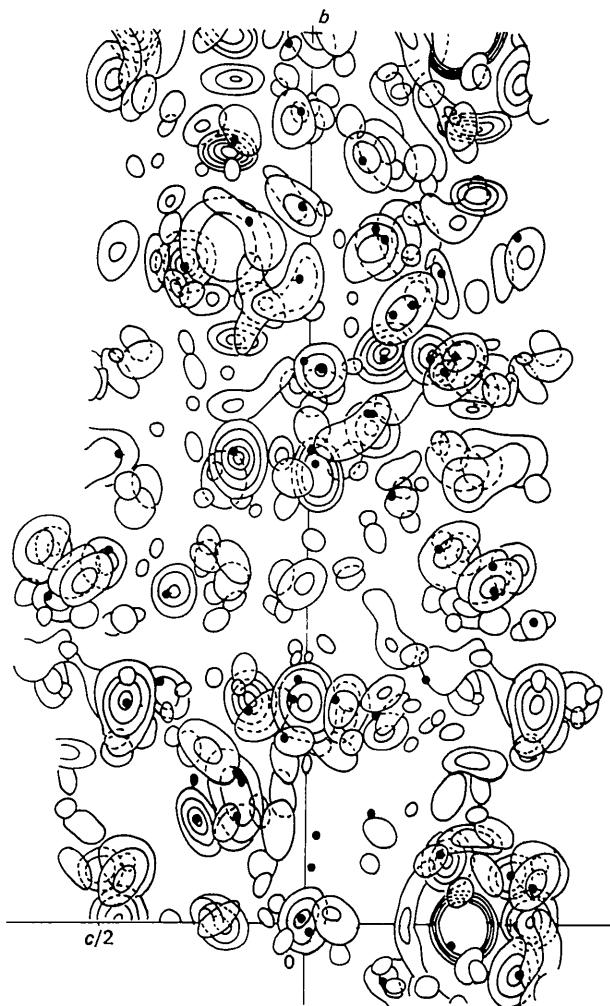


Fig. 1. The composite Fourier map projected along the a axis. Phases are based on two iodine atoms. Contour interval is arbitrary. Black circlets indicate the final atomic positions of two molecules in the asymmetric unit.

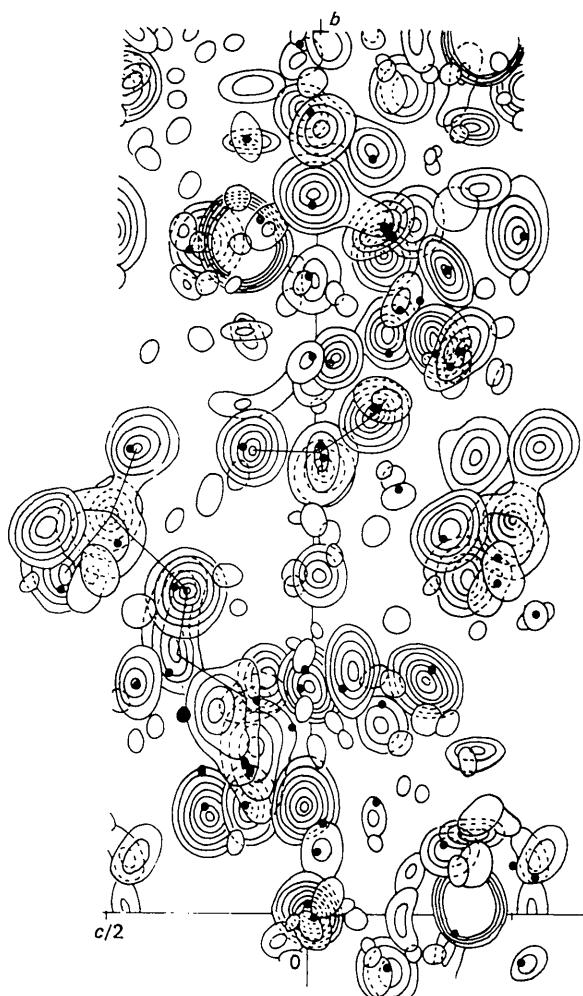


Fig. 2. The composite Fourier map projected along the a axis. Phases are based on two iodine and thirty-six light atoms. Contour interval is arbitrary. Black circlets indicate the final atomic positions of two molecules.

ranging between 0.22 and 0.42 were included. Three more cycles of refinement were made, using all the reflexions. The final *R* value dropped to 0.157. The final atomic coordinates and temperature factors for each atom are given in Table 1, and the observed and

Table 1. The final atomic coordinates and temperature factors (\AA^2)

For the iodine atoms, anisotropic temperature factors in the form of $\exp \{-(h^2B_{11}+k^2B_{22}+l^2B_{33}+hkB_{12}+klB_{23}+lhB_{31})\}$ were applied.

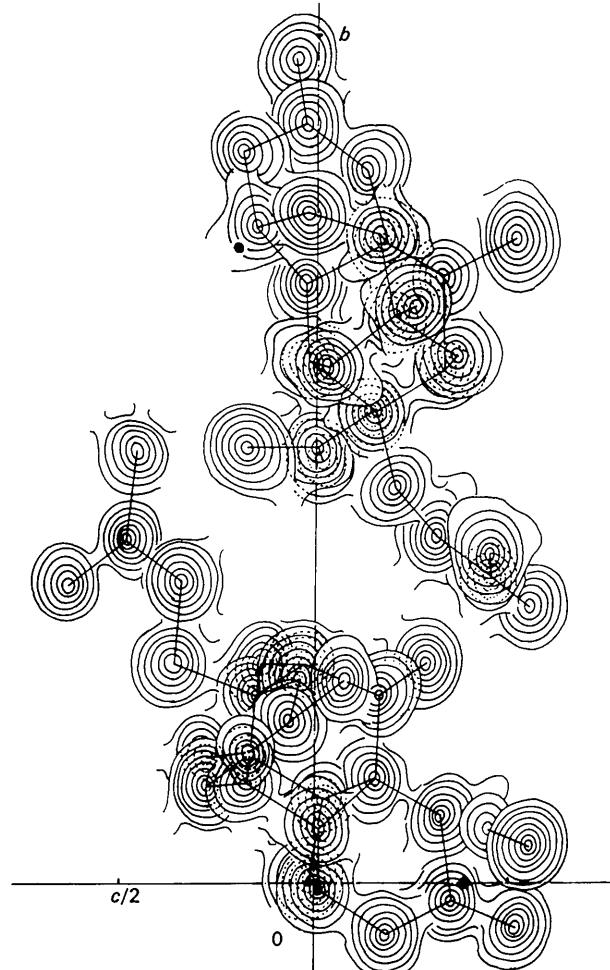


Fig. 3. The final composite Fourier map projected along the *a* axis, contours being drawn at intervals of $1 \text{ e.}\text{\AA}^{-3}$. Contours for the iodine atoms are omitted and their positions are shown with black circles.

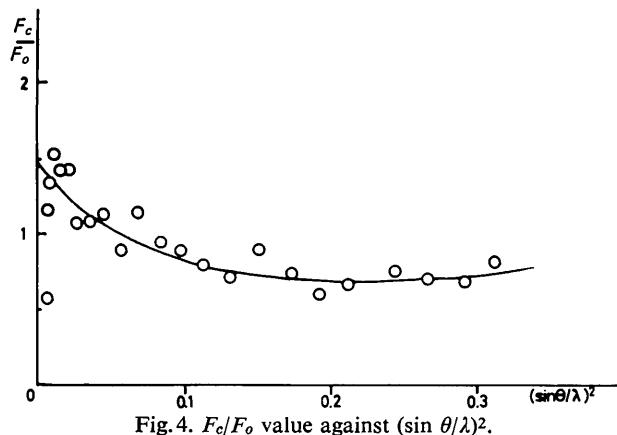


Fig. 4. F_c/F_o value against $(\sin \theta/\lambda)^2$.

Atom	<i>x/a</i>	<i>y/b</i>	<i>z/c</i>	<i>B</i>
N(1)	0.355	0.151	-0.122	1.2
C(2)	0.362	0.107	0.033	0.7
N(3)	0.406	0.137	0.182	2.0
C(4)	0.445	0.221	0.155	1.4
C(5)	0.418	0.376	0.013	2.5
C(6)	0.377	0.437	-0.157	0.7
C(7)	0.381	0.375	-0.329	1.5
C(8)	0.346	0.280	-0.327	1.5
C(9)	0.387	0.235	-0.156	1.3
C(10)	0.392	0.284	0.029	0.9
N(11)	0.332	0.025	0.056	2.8
O(12)	0.492	0.378	-0.033	1.5
O(13)	0.304	0.436	-0.141	1.9
C(14)	0.413	0.524	-0.182	3.1
O(15)	0.364	0.585	-0.311	2.7
O(16)	0.454	0.365	-0.354	1.9
O(17)	0.346	0.224	-0.493	1.6
O(18)	0.493	0.203	0.024	1.7
C(19)	0.465	0.231	-0.160	1.6
C(20)	0.495	0.319	-0.204	1.2
O(21)	0.561	0.313	-0.237	3.4
C(22)	0.287	0.491	0.006	2.6
O(23)	0.321	0.495	0.165	3.0
C(24)	0.203	0.501	-0.020	3.2
C(25)	0.395	0.621	-0.445	1.6
C(26)	0.343	0.674	-0.559	2.9
O(27)	0.457	0.609	-0.464	3.5
N(51)	0.013	0.931	-0.329	0.5
C(52)	0.026	1.015	-0.355	1.7
N(53)	0.051	1.069	-0.198	0.9
C(54)	0.055	1.020	-0.005	1.0
C(55)	0.114	0.883	0.158	0.0
C(56)	0.128	0.776	0.146	0.7
C(57)	0.061	0.736	0.025	1.0
C(58)	0.042	0.775	-0.168	0.0
C(59)	0.029	0.887	-0.152	0.0
C(60)	0.090	0.932	-0.028	0.7
N(61)	0.011	1.055	-0.523	1.8
O(62)	0.052	0.887	0.264	0.7
O(63)	0.180	0.761	0.016	2.1
C(64)	0.146	0.736	0.338	3.0
O(65)	0.157	0.642	0.329	1.0
O(66)	0.001	0.756	0.134	0.9
O(67)	-0.013	0.736	-0.274	1.1
O(68)	-0.014	0.998	0.026	0.4
C(69)	-0.026	0.906	-0.019	0.8
C(70)	-0.009	0.852	0.175	2.0
O(71)	-0.057	0.860	0.295	1.2
C(72)	0.241	0.804	0.058	2.2
O(73)	0.267	0.858	0.166	3.8
C(74)	0.287	0.761	-0.086	2.6
C(75)	0.162	0.598	0.483	0.7
C(76)	0.165	0.497	0.462	2.5
O(77)	0.162	0.644	0.634	2.0
O(78)	0.390	0.961	-0.553	3.2
C(79)	0.433	0.949	-0.376	3.5
I(81)	0.2253	0.0000	-0.3941	*
I(82)	0.1902	0.2433	0.1903	†

B_{11} B_{22} B_{33} B_{12} B_{23} $B_{31} (\times 10^5)$

* I(81) 121 123 543 -50 -306 51

† I(82) 94 116 1195 97 296 13

Table 2. Observed and calculated structure factors ($\times 10$)

<i>h</i>	<i>Po</i>	<i>Pc</i>	<i>b</i>	<i>Po</i>	<i>Pc</i>	<i>h</i>	<i>Po</i>	<i>Pc</i>																					
0	6	7	215	215	15	333	340	2	301	312	8	512	505	15	143	144	13	216	242	10	243	313	1	718	773				
1	123	150	9	137	169	1	334	300	4	240	225	10	578	496	17	218	220	15	219	165	12	176	158	3	507	453			
2	112	106	10	580	630	18	208	191	5	225	220	11	398	351	18	0	32	16	266	289	13	144	108	0	133	96			
3	117	103	11	215	219	19	178	117	6	229	236	12	366	316	19	0	181	17	0	48	16	12	132	1	133	83			
4	5	6	334	376	12	376	12	20	268	270	7	179	122	20	127	104	8	230	210	15	176	124	6	366	400				
5	0	50	11	338	521	21	141	143	8	183	201	14	267	277	11	139	19	8	81	51	17	141	166	3	230	211			
6	114	41	16	271	270	22	233	198	9	193	200	15	181	161	22	107	134	7	242	318	17	251	291	5	133	173			
7	96	11	15	410	478	23	111	128	10	138	99	218	187		2	4	18	250	291	5	133	173	9	318	304				
8	0	11	16	178	166	4	233	193	11	189	218	17	207	330	1	-6	158	196	6	175	70	10	249	207					
9	82	121	17	333	345	0	-5	12	247	265	14	208	195	202	0	266	290	2	169	165	21	97	67	8	225	280			
0	7	19	199	130	0	547	631	14	208	200	20	186	178	1	182	180	0	466	422	2	247	282	9	161	161				
1	430	346	21	0	0	2	536	527	16	161	147	22	153	128	3	195	171	5	165	190	26	192	125	11	0	32			
2	633	459	24	111	116	16	226	144	14	482	502	24	115	102	6	288	346	8	216	225	12	183	213	18	156	109			
3	0	24	0	90	0	1	351	365	1	-1	7	190	203	9	72	75	0	780	678	16	173	170	19	126	112				
4	210	237	0	0	7	295	294	2	325	300	8	138	136	10	72	68	1	395	344	16	173	213	18	156	109				
5	0	15	8	155	162	3	153	156	0	221	50	9	119	121	11	156	152	2	305	262	17	234	244	21	127	85			
6	157	239	2	158	5	301	311	4	345	355	1	189	87	10	199	222	12	126	126	3	595	560							
7	116	126	3	147	25	10	300	341	5	278	285	2	394	304	11	186	173	13	13	71	4	468	379	2	-8	3			
8	102	65	4	395	409	11	157	183	6	280	285	3	571	615	12	146	105	14	104	98	5	63	810						
9	11	70	5	944	1206	12	162	159	7	285	308	4	699	735	13	204	251	15	274	265	6	353	321	0	216	212			
10	12	0	3	553	553	13	134	71	8	282	315	5	641	765	16	123	161	7	120	120	2	246	249	1	150	106			
11	57	91	7	411	393	14	145	147	6	186	147	10	300	357	15	200	187	10	88	142	2	213	231	3	194	110			
12	0	88	8	184	123	15	149	147	9	226	226	8	436	459	17	181	154	19	126	126	10	220	204	4	482	523			
13	0	15	17	19	178	16	142	166	22	116	122	17	236	236	19	193	187	18	207	284	11	125	223	5	0	57			
14	0	10	190	173	17	0	31	12	219	236	23	367	367	18	199	182	20	278	284	11	125	521	6	181	146				
15	124	68	11	0	20	18	0	45	121	211	10	633	732	19	143	107	13	124	124	12	146	146	7	555	588				
16	245	281	1	17	19	19	0	81	120	208	12	179	179	17	190	190	17	20	212	177	2	3	12	12	12				
17	0	15	24	1	0	7	120	120	11	146	146	5	224	225	17	146	146	12	146	146	12	146	146	5	250	269			
18	4	173	149	95	21	95	86	165	165	13	397	505	1	-7	1	532	606	15	465	415	9	0	15	10	205	204			
19	5	267	269	15	88	32	22	124	183	17	207	225	14	210	190	14	188	188	21	124	124	10	0	64	11	11	237		
20	6	206	16	177	85	17	0	58	94	58	31	370	0	280	281	3	304	303	17	503	517	10	0	54	12	515	442		
21	0	17	17	17	0	6	0	6	10	224	216	14	257	257	5	836	757	19	197	157	2	295	261	14	150	138			
22	0	22	19	181	190	0	382	395	1	333	472	19	251	256	4	288	214	7	593	488	21	0	77	7	300	268			
23	10	255	328	20	93	18	13	139	61	1	333	472	19	251	256	4	288	214	7	593	488	21	0	77	7	180	164		
24	11	167	167	21	125	104	2	129	59	2	441	436	20	134	137	5	297	228	8	690	690	21	0	77	7	180	164		
25	12	215	224	22	125	11	151	151	6	414	421	22	171	176	7	187	182	18	459	509	2	-3	2	213	223	19	0	56	
26	14	178	184	24	133	5	151	132	5	379	422	23	151	186	8	305	263	11	191	146	19	193	207	3	199	217			
27	15	293	275	0	-1	8	301	316	24	168	216	21	241	272	1	257	260	21	53	41	9	141	103	3	157	79			
28	0	5	0	461	436	9	331	378	9	205	168	6	415	442	12	139	162	6	257	260	15	409	353	8	0	149	5	484	506
29	1	131	150	1	150	10	273	312	0	588	570	13	203	180	16	136	136	7	409	378	15	1	746	810					
30	2	185	167	10	252	198	11	175	242	1	410	347	14	252	240	17	164	157	5	207	230	15	404	395	3	184	167		
31	3	515	541	3	541	541	1	177	240	12	135	82	2	380	315	17	175	187	10	245	245	11	365	353					
32	5	575	575	3	575	575	1	177	240	12	135	82	2	380	315	17	175	187	10	245	245	11	365	353					
33	6	526	526	3	526	526	0	135	4	567	592	14	265	303	2	210	153	8	245	292	19	0	245	3	322	311			
34	7	578	508	6	508	6	204	249	5	344	366	26	17	267	253	2	210	153	8	245	292	19	0	245	3	322	311		
35	8	563	563	6	563	6	204	249	5	344	366	26	17	267	253	2	210	153	8	245	292	19	0	245	3	322	311		
36	9	519	519	7	519	519	0	135	4	567	592	14	265	303	2	210	153	8	245	292	19	0	245	3	322	311			
37	10	555	555	8	555	555	0	135	4	567	592	14	265	303	2	210	153	8	245	292	19	0	245	3	322	311			
38	11	522	522	9	522	522	0	135	4	567	592	14	265	303	2	210	153	8	245	292	19	0	245	3	322	311			
39	12	522	522	10	522	522	0	135	4	567	592	14	265	303	2	210	153	8	245	292	19	0	245	3	322	311			
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41	14	522	522	12	522	522	2	5	522	522	14	265	303	2	210	153	8	245	292	19	0	245	3	322	311				
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Table 2 (cont.)

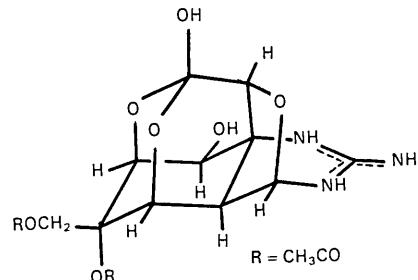
b.	Po	Pc	h	Po	Pc	b.	Po	Pc	h	Po	Pc	b.	Po	Pc	h	Po	Pc															
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16	288	296	10	114	98	18	93	147	8	191	163	11	138	155	23	121	105	11	164	163	15	192	243	10	241	161	1					
17	330	320	11	161	132	19	55	60	8	193	235	12	154	156	5	—	—	12	0	125	17	120	134	11	319	286	2					
18	165	140	11	154	166	20	90	111	9	92	85	13	139	211	16	193	219	14	133	26	18	0	43	12	0	97	3					
19	173	121	13	128	73	21	71	53	10	192	221	16	193	219	19	173	219	14	133	26	18	0	43	12	0	97	4					
20	129	16	14	152	161	4	1	—	—	12	233	276	15	0	58	0	612	629	15	206	263	19	170	147	21	0	39					
21	117	112	—	—	—	—	—	—	1	233	276	12	289	302	18	92	97	17	170	147	21	0	49	15	0	15	0					
22	150	206	3	—	9	13	24	112	5	5	—	—	2	315	268	17	170	147	6	2	17	332	258	8	173	172	7	0				
23	74	62	1	119	141	2	125	41	15	267	308	1	366	309	9	381	370	19	135	141	17	312	233	20	315	258	10	157	101			
24	108	168	0	119	141	3	124	34	15	267	308	5	370	370	5	—	—	2	244	199	20	315	258	8	173	172	11	0				
3	—	3	2	117	121	2	124	112	5	5	—	—	2	324	258	5	—	—	3	213	131	21	304	273	12	161	97	7	5			
0	431	361	4	136	70	6	257	226	19	242	172	5	320	246	8	411	445	0	250	266	3	213	131	22	183	148	13	136	117			
1	513	513	5	168	193	7	286	217	20	287	250	6	342	298	9	390	369	1	208	203	6	170	63	17	23	0	116	116				
2	400	400	6	148	134	9	477	428	28	301	260	7	239	199	10	360	340	12	205	187	6	196	140	15	204	187	11	116	104			
3	376	267	8	134	114	9	114	63	2	278	262	8	147	153	11	350	373	3	172	199	7	237	142	6	4	4	4	4				
4	620	566	8	0	150	11	498	457	9	182	166	10	412	400	4	374	322	6	196	121	1	246	229	8	173	172	11	0				
5	361	343	11	259	255	4	—	—	10	210	250	13	402	426	5	464	406	12	217	201	6	216	0	358	304	2	170	110				
6	131	113	4	8	12	558	450	11	116	156	12	364	342	12	367	340	7	145	151	10	217	201	6	216	0	358	304	2	170	110		
7	371	371	12	120	154	14	182	135	1	207	178	13	0	26	16	210	167	8	116	116	12	233	216	3	252	204	5	239	218			
8	122	130	1	120	154	16	182	135	1	207	178	13	0	26	16	210	167	8	116	116	12	233	216	3	252	204	5	239	218			
9	336	366	2	84	65	15	570	558	2	649	669	16	135	122	17	294	266	13	191	199	13	191	199	13	191	199	13	191	199			
10	263	270	3	74	76	16	130	73	3	367	409	15	161	152	18	128	130	11	151	183	16	405	372	5	297	265	7	294	350			
11	273	273	4	68	87	17	383	409	4	222	211	16	155	126	12	201	194	11	151	183	4	757	668	19	84	59	1	198	208			
12	240	255	5	94	117	17	383	409	5	165	153	17	130	145	20	216	224	12	144	114	15	196	162	19	84	59	1	198	208			
13	240	255	6	87	85	19	163	155	6	136	96	21	0	147	13	138	132	17	144	114	15	196	162	19	84	59	1	198	208			
14	213	194	7	101	134	20	244	306	7	562	540	5	6	22	113	101	14	174	186	18	180	69	49	9	306	410	11	196	72			
15	438	390	8	0	34	21	75	59	8	227	227	13	170	170	17	170	170	12	191	191	21	20	0	36	113	13	196	72				
16	287	246	9	0	77	22	152	221	9	79	77	13	1	353	332	5	—	—	17	90	40	21	273	301	12	0	134	180	187			
17	0	77	4	7	23	102	150	10	0	108	108	2	367	341	5	—	—	17	90	40	21	273	301	13	0	124	182	187				
18	246	206	11	141	160	13	271	320	4	343	277	0	500	565	5	—	—	17	90	40	21	273	301	13	0	153	160	160				
19	95	68	1	314	329	4	0	—	13	0	117	5	248	268	1	376	351	0	144	146	1	378	322	15	160	146	1	272	212			
20	211	223	2	0	92	13	299	221	2	618	866	16	140	133	2	337	316	0	144	146	1	378	322	15	160	146	1	272	212			
21	178	167	3	259	221	2	618	866	16	140	133	2	337	316	0	144	146	1	378	322	15	160	146	1	272	212						
22	108	101	2	257	229	3	618	866	16	140	133	2	337	316	0	144	146	1	378	322	15	160	146	1	272	212						
23	96	88	5	166	163	6	459	533	16	0	13	9	247	226	5	294	280	3	211	165	17	235	205	17	0	40	3	278	250			
3	—	4	7	105	83	6	376	262	18	60	62	10	173	161	12	304	282	1	204	199	12	235	215	17	0	40	3	278	250			
4	193	239	8	193	239	7	502	452	16	0	157	119	1	204	199	10	204	199	16	160	140	17	204	199	16	160	140	17	204	199		
5	521	490	11	115	104	12	81	70	4	—	—	15	157	119	1	204	199	16	160	140	17	204	199	16	160	140	17	204	199			
6	261	221	12	124	121	13	165	93	5	165	68	17	170	170	17	170	170	13	165	165	17	205	185	13	165	165	17	205	185			
7	234	214	13	124	121	14	165	68	1	0	65	1	452	574	19	233	293	1	0	158	7	180	119	5	197	122	13	165	165			
8	223	201	14	124	121	15	227	257	6	0	65	1	452	574	19	233	293	1	0	158	7	180	119	5	197	122	13	165	165			
9	237	219	15	124	121	16	242	266	0	163	152	10	399	401	11	201	213	10	225	235	1	191	191	16	201	161	13	237	219			
10	240	228	13	124	121	17	30	93	8	171	174	7	161	213	2	328	226	12	172	172	1	301	161	13	237	219	16	201	161			
11	246	226	14	124	121	18	30	93	8	171	174	7	161	213	2	328	226	12	172	172	1	301	161	13	237	219	16	201	161			
12	245	225	15	124	121	19	30	56	10	170	177	3	500	522	20	153	166	5	0	502	522	20	153	166	5	0	502	522				
13	161	162	16	71	91	21	187	229	11	107	84	9	170	129	21	131	131	6	176	176	11	232	212	11	0	109	7	232	212			
14	236	263	17	116	188	12	228	228	83	12	97	108	7	425	457	22	132	132	7	187	187	11	232	212	11	0	109	7	232	212		
15	181	181	8	0	35	13	86	93	10	143	153	11	311	311	5	—	—	9	0	55	6	—	1	235	221	8	235	221	8	235	221	
16	179	188	9	4	29	7	26	143	10	143	153	5	274	274	5	—	—	10	0	55	6	—	1	235	221	8	235	221	8	235	221	
17	0	36	1	197	167	4	—	—	16	60	57	11	285	265	5	—	—	11	197	167	4	—	—	1	235	221	8	235	221	8	235	221
18	201	188	2	34	321	32	103	103	17	60	57	11	28																			

Table 2 (cont.)

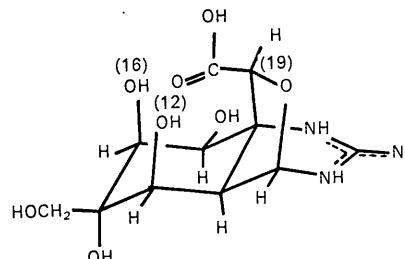
b	Po	Pc	n	Po	Pc	b	Po	Pc	n	Po	Pc	b	Po	Pc	n	Po	Pc	b	Po	Pc	n	Po	Pc	b	Po	Pc
---	----	----	---	----	----	---	----	----	---	----	----	---	----	----	---	----	----	---	----	----	---	----	----	---	----	----

Table 2 (cont.)

<i>h</i>	<i>Po</i>	<i>Fc</i>																																		
12	2		16	121	103	14	150	119	11	130	73	10	0	132	4	212	204	16	168	154	12	209	226	9	0	181										
1	225	206	12	1		16	0	17	13	60	95	12	253	258	6	187	1	187	17	134	78	13	220	244	10	0	170									
2	201	177				17	231	246	14	21	113	13	133	3	122	2	154	2	154	10	18	14	16	0	86	11	142	62								
3	450	512	1	181	186				12	-1		14	336	295	9	173	184	3	389	390	12	-4	15	238	202	13	0	154								
4	284	269	2	142	71	12	0			16	270	275	10	0	81	9	151	113	4	293	276	16	172	61	13	0	62									
5	262	237	3	210	178				1	180	298	0	121	70	17	190	161	11	164	161	0	112	143	12	-5		14	0	59							
6	446		4	151	149				1	180	298	0	121	70	17	190	161	11	164	161	0	112	143	12	-5		12	-6								
7	135	25	5	167	175	2	312	471	1	187	155	18	146	154	12	113	192	13	130	121	8	303	386	6	336	425	1	0	62							
8	201	72	6	189	163	3	280	283	2	233	230	19	165	192	14	138	126	9	210	213	16	209	111	0	319	340	1	0	71							
9	293	287	7	217	225	4	317	298	3	186	153	18	153	186	15	124	195	10	192	250	5	143	45	3	192	207	1	0	109							
10	277	313	8	293	200	5	428	432	4	185	250	12	-2		16	149	230	11	159	265	6	0	69	4	240	179	3	0	125							
11	24	29	9	181	138	6	237	117	7	187	157	18	157	187	17	0	56	1	0	43	7	122	31	5	265	283	4	0	90							
12	0	71	10	269	276	7	263	261	6	156	213	0	251	196	17	0	56	1	0	43	149	193	18	74	14	139	113	9	216	201	6	222	54	5	0	111
13	191	172	11	264	227	8	152	175	7	306	453	1	335	303	18	74	149	193	18	198	240	14	139	113	9	109	187	7	277	318	6	0	61			
14	0	75	12	269	309	9	235	233	8	0	75	2	224	236	219	12	-3	15	156	184	11	134	161	8	215	167										
15	168	193	13	187	209	10	176	180	9	291	359	3	235	219																						



(IV)



(V)

calculated structure factors are listed in Table 2. Computations were done on an IBM 7090 computer with programs written by the authors.

Discussion

Fig. 3 shows the superimposed sections of the electron density distribution cut through the atomic centres of DATI and methyl alcohol. It is conceived that the general structure of the DAT molecule is now established, except its absolute configuration. The fact that essentially the same structure has been deduced from the two independent molecules in the asymmetric unit is additional evidence of the reliability of the present analysis. In the present paper, except in Figs. 1, 2 and 3, the molecules are drawn with the correct absolute configuration, by taking account of the result by Tomie, Furusaki, Kasami, Yasuoka, Miyake, Haisa & Nitta (1964) who have determined the absolute configuration of bromoanhydrotetrodionic lactone hydrobromide, another derivative of tetrodotoxin.

The present compound and tetrodonic acid have the same perhydroquinazoline skeleton. The conformation of substituent atoms to the cyclohexane ring is the same for these two compounds; O(17) and C(14) are in equatorial positions and O(13) is in axial position. The structural relationship between DAT and tetrodonic acid may be described as follows: Let us start from the hypothetical epitetrodonic acid (V), in which the configuration at C(19) is antipodal to that in tetrodonic acid. The ortho-ester may, then, be formed by combining the acid carbonyl with the hydroxyl groups, O(12) and O(16). The skeletal structure of DAT so built up is similar to that of adamantine, which is a cage molecule consisting of four six-membered rings in the chair form.

Intramolecular bond lengths and angles in the two independent molecules are shown in Fig. 5 and Table 3 respectively. The estimated standard deviations in the atomic coordinates are approximately 0.004 Å for the iodine and 0.07 Å for the light atoms. Hence, the standard deviation of the bond lengths between the light

Table 3. Bond angles

C(2)—N(1)—C(9)	126°	C(52)—N(51)—C(59)	123°
N(1)—C(2)—N(3)	121	N(51)—C(52)—N(61)	119
N(1)—C(2)—N(11)	124	N(53)—C(52)—N(61)	119
N(3)—C(2)—N(11)	115	C(52)—N(53)—C(54)	115
C(2)—N(3)—C(4)	117	C(53)—C(54)—C(60)	106
N(3)—C(4)—C(10)	105	N(53)—C(54)—O(68)	108
N(3)—C(4)—O(18)	109	C(60)—C(54)—O(68)	106
C(10)—C(4)—O(18)	99	C(56)—C(55)—C(60)	117
C(6)—C(5)—C(10)	115	C(56)—C(55)—O(62)	103
C(6)—C(5)—O(12)	102	C(60)—C(55)—O(62)	105
C(10)—C(5)—O(12)	113	C(55)—C(56)—C(57)	105
C(5)—C(6)—C(7)	102	C(55)—C(56)—O(63)	108
C(5)—C(6)—O(13)	109	C(55)—C(56)—C(64)	116
C(5)—C(6)—C(14)	112	C(57)—C(56)—C(64)	115
C(7)—C(6)—O(13)	102	C(57)—C(56)—O(63)	100
C(7)—C(6)—C(14)	110	C(57)—C(56)—C(64)	116
O(13)—C(6)—C(14)	120	O(63)—C(56)—C(64)	115
C(6)—C(7)—C(8)	117	C(56)—C(57)—C(58)	117
C(6)—C(7)—O(16)	108	C(56)—C(57)—O(66)	110
C(8)—C(7)—O(16)	110	C(58)—C(57)—O(66)	107
C(7)—C(8)—C(9)	102	C(57)—C(58)—C(59)	109
C(7)—C(8)—O(17)	117	C(57)—C(58)—O(67)	116
C(9)—C(8)—O(17)	111	C(59)—C(58)—O(67)	110
N(1)—C(9)—C(8)	109	N(51)—C(59)—C(58)	114
N(1)—C(9)—C(10)	104	N(51)—C(59)—C(60)	113
N(1)—C(9)—C(19)	115	N(51)—C(59)—C(69)	113
C(8)—C(9)—C(10)	118	C(58)—C(59)—C(60)	111
C(8)—C(9)—C(19)	115	C(58)—C(59)—C(69)	110
C(10)—C(9)—C(19)	95	C(60)—C(59)—C(69)	96
C(4)—C(10)—C(5)	113	C(54)—C(60)—C(55)	114
C(4)—C(10)—C(9)	101	C(54)—C(60)—C(59)	96
C(5)—C(10)—C(9)	111	C(55)—C(60)—C(59)	115
C(5)—O(12)—C(20)	108	C(55)—O(62)—C(70)	116
C(6)—O(13)—C(22)	113	C(56)—O(63)—C(72)	116

Table 3 (cont.)

C(6)—C(14)—O(15)	108	C(56)—C(64)—O(65)	111
C(14)—O(15)—C(25)	112	C(64)—O(65)—C(75)	117
C(7)—O(16)—C(20)	114	C(57)—O(66)—C(70)	115
C(4)—O(18)—C(19)	111	C(54)—O(68)—C(69)	108
C(9)—C(19)—O(18)	105	C(59)—C(69)—O(68)	102
C(9)—C(19)—C(20)	113	C(59)—C(69)—C(70)	112
O(18)—C(19)—C(20)	110	O(68)—C(69)—C(70)	105
O(12)—C(20)—O(16)	106	O(62)—C(70)—O(66)	109
O(12)—C(20)—C(19)	106	O(62)—C(70)—C(69)	107
O(12)—C(20)—O(21)	109	O(62)—C(70)—O(71)	107
O(16)—C(20)—C(19)	112	O(66)—C(70)—C(69)	109
O(16)—C(20)—O(21)	111	O(66)—C(70)—O(71)	109
C(19)—C(20)—O(21)	113	C(69)—C(70)—O(71)	116
O(13)—C(22)—O(23)	125	O(63)—C(72)—O(73)	139
O(13)—C(22)—C(24)	107	O(63)—C(72)—C(74)	103
O(23)—C(22)—C(24)	120	O(73)—C(72)—C(74)	118
O(15)—C(25)—C(26)	106	O(65)—C(75)—C(76)	114
O(15)—C(25)—O(27)	125	O(65)—C(75)—O(77)	118
C(26)—C(25)—O(27)	129	C(76)—C(75)—O(77)	128

atoms is about 0.1 Å, and that of the bond angles is about 6°. None of the bond lengths and angles deviate significantly from their normal values. The average values of the bond lengths are 1.53 Å for C—C, 1.33 Å for C—N (in the guanidine group), 1.46 Å for C—N, 1.42 Å for C—O and 1.23 Å for C=O. It should be pointed out that the corresponding bond lengths and angles in the two independent molecules are in good agreement as seen from Fig. 5 and Table 3.

The crystal structure viewed along the *c* axis is illustrated in Fig. 6, where anticipated hydrogen bonds are shown by the broken lines. The hydrogen bond lengths and short intermolecular contacts are listed in Table 4. The donors of the hydrogen bonds in Table 4 are the atoms at *x*, *y*, *z*. There are six NH···O, three OH···N and one OH···O hydrogen bond per asymmetric unit, and these hydrogen bonds connect the

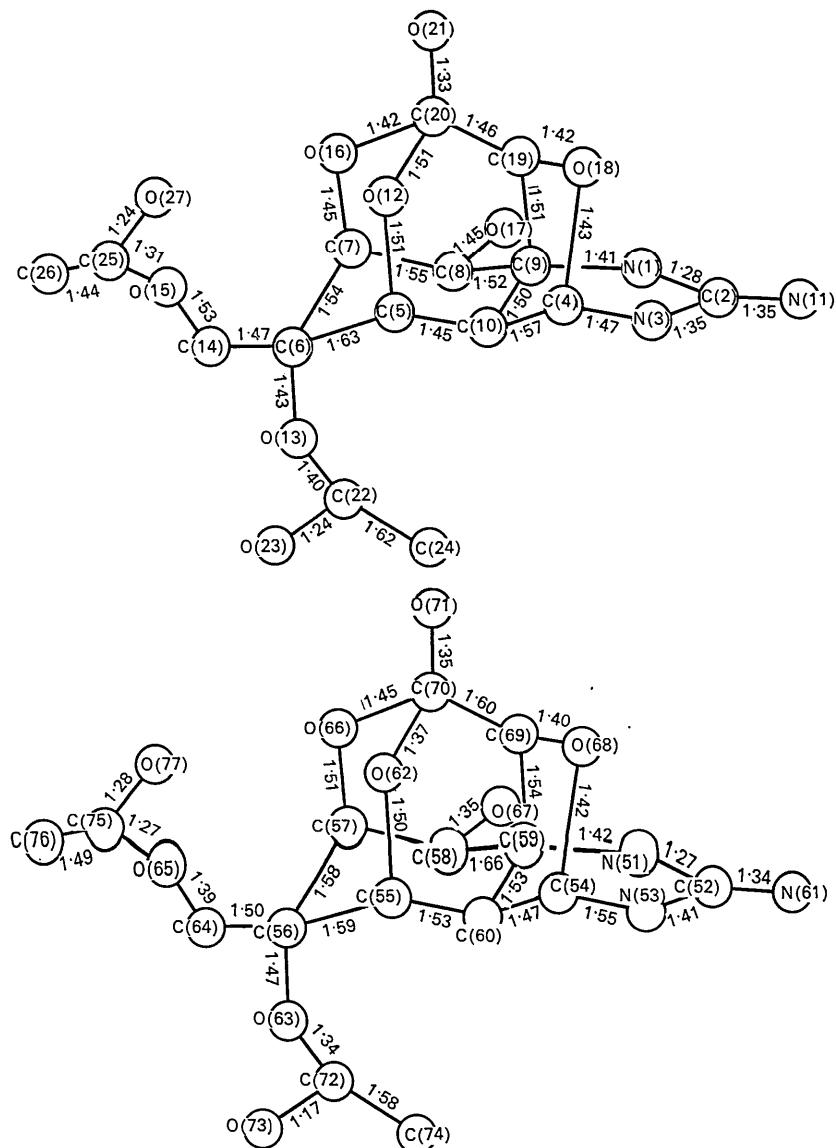


Fig. 5. Bond lengths (Å).

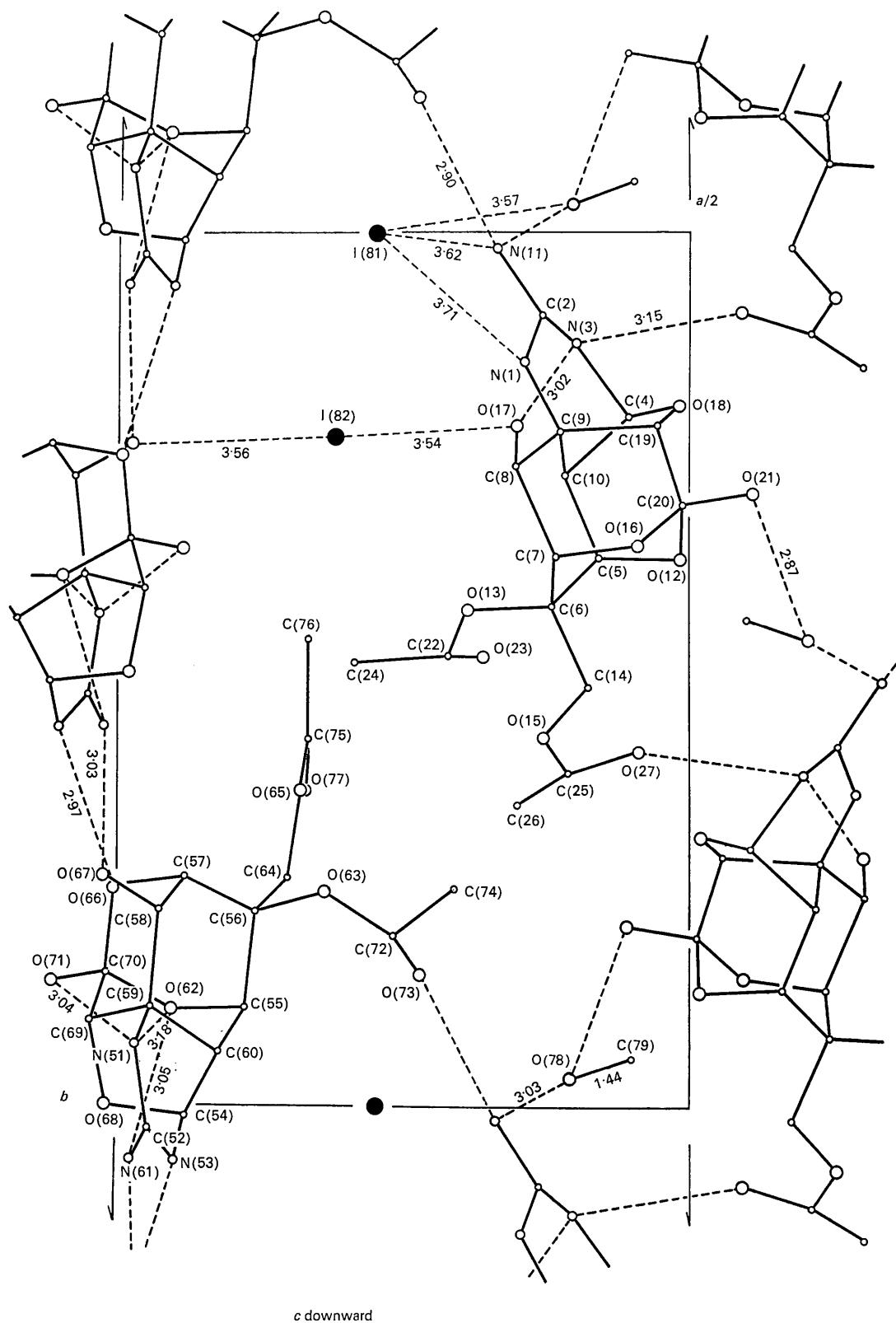


Fig. 6. Crystal structure viewed along the *c* axis.

DAT and methyl alcohol molecules to form an infinite three-dimensional network. The iodine atoms are in the interstices of the three-dimensional network and short contacts are found for I(81)· · · N(1), I(81)· · · N(11), I(82)· · · O(17) and I(82)· · · O(67). It may be suggested that the iodine atoms interact with the hydrogen atoms bonded with these nitrogen or oxygen atoms.

Table 4. Possible hydrogen bonds and short intermolecular distances

From atom at <i>x, y, z</i>	To atom	At position	Distance
--------------------------------	------------	----------------	----------

(a) Possible hydrogen bonds

N(3)	O(27)	11	3.15 Å
N(11)	O(73)	4	2.90
O(17)	N(3)	2	3.02
O(21)	O(78)	12	2.87
N(51)	O(62)	2	3.18
N(53)	O(66)	9	2.97
N(61)	O(62)	2	3.05
N(61)	O(67)	10	3.03
O(71)	N(51)	3	3.04
O(78)	N(11)	5	3.03

(b) Short intermolecular contacts

I(81)	N(1)	1	3.71
I(81)	N(11)	1	3.62
I(81)	O(78)	4	3.57
I(81)	N(53)	4	3.95
I(81)	C(52)	4	3.91
I(81)	O(73)	6	3.95
I(82)	O(17)	3	3.54
I(82)	O(67)	7	3.56
O(15)	C(74)	1	3.48
C(24)	O(65)	1	3.45
C(24)	C(76)	1	3.63
C(26)	C(74)	1	3.91
O(17)	C(4)	2	3.36
C(24)	C(75)	2	3.83
C(24)	C(76)	2	3.70
C(24)	O(77)	2	3.27
C(25)	O(23)	2	3.50
C(26)	O(23)	2	3.28
C(26)	C(64)	2	3.89
C(26)	C(72)	2	3.70
C(26)	O(73)	2	3.54
C(26)	C(75)	2	3.72
C(52)	O(62)	2	3.41
C(58)	O(77)	2	3.45
N(61)	O(68)	2	3.32
N(61)	O(71)	2	3.34
O(63)	O(77)	2	3.21
O(67)	O(71)	2	3.59
C(74)	O(77)	2	3.40
O(78)	O(73)	2	3.28
O(78)	N(3)	5	3.24
C(24)	C(69)	7	3.74
C(57)	N(53)	7	3.59
C(57)	C(54)	7	3.87
C(75)	C(52)	7	3.83
C(75)	N(61)	7	3.45
C(76)	N(51)	7	3.58
C(76)	C(52)	7	3.68
C(76)	N(61)	7	3.59

Table 4 (cont.)

From atom at <i>x, y, z</i>	To atom	At position	Distance
O(77)	N(61)	7	3.58
C(76)	O(71)	8	3.52
C(4)	C(14)	11	3.97
C(4)	C(25)	11	3.79
C(4)	O(27)	11	3.18
C(5)	C(79)	11	3.78
O(12)	C(79)	11	3.27
O(18)	C(14)	11	3.30
O(18)	C(27)	11	3.47
O(21)	O(73)	11	3.36
O(21)	C(74)	11	3.58
O(16)	O(78)	12	3.47
O(16)	C(79)	12	3.34
C(20)	C(79)	12	3.97
O(21)	C(26)	12	3.23
O(21)	C(79)	12	3.43
O(27)	C(79)	12	3.46
Position 1	<i>x</i>	<i>y</i>	<i>z</i>
2	<i>x</i>	<i>y</i> - 1 + <i>z</i>	8 - <i>x</i> - $\frac{1}{2}$ + <i>y</i> - <i>z</i>
3	<i>x</i>	<i>y</i> 1 + <i>z</i>	9 - <i>x</i> $\frac{1}{2}$ + <i>y</i> - <i>z</i>
4	<i>x</i> - 1 + <i>y</i>	<i>z</i>	10 - <i>x</i> $\frac{1}{2}$ + <i>y</i> - 1 - <i>z</i>
5	<i>x</i> 1 + <i>y</i> - 1 + <i>z</i>	11 1 - <i>x</i> - $\frac{1}{2}$ + <i>y</i> - <i>z</i>	
6	<i>x</i> - 1 + <i>y</i> - 1 + <i>z</i>	12 1 - <i>x</i> - $\frac{1}{2}$ + <i>y</i> - 1 - <i>z</i>	

All distances less than 4.0 Å between the reference molecules at *x, y, z* and the neighbouring molecules were calculated. Only those less than 4.0 Å for C · · · C and C · · · I pairs and those less than 3.6 Å for the other pairs are listed.

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